DR. BABASAHEB AMBEDKAR MARATHWADA UNIVERSITY

CIRCULAR NO.SU/Sci./M.Sc.Chemistry/59/2021

It is hereby inform to all concerned that, the syllabus prepared by the Board of Studies in Mathematics and recommended by the Dean, Faculty of Science & Technology the Hon'ble Vice-Chancellor has accepted the **Syllabus of M.Sc.Chemistry Ist to IVth semester with Bridge Course for affiliated Colleges and University Department** in his emergency powers under section 12(7) of the Maharashtra Public Universities Act, 2016 on behalf of the Academic Council as appended herewith.

This shall be effective from the Academic Year 2021-22 and onwards.

All concerned are requested to note the contents of this circular and bring notice to the students, teachers and staff for their information and necessary action.

University Campus,
Aurangabad-431 004.
REF.NO. SU/SCI/2021/4199-208
Date:- 29-11-2021.

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Deputy Registrar,
Academic Section.

Copy forwarded with compliments to :-

1] **The Principal of all concerned Colleges,**
Dr. Babasaheb Ambedkar Marathwada University,
2] **Head of the Department, Department of Microbiology,**
Dr.Babasaheb Ambedkar Marathwada University, Aurangabad.
3] **The Director, University Network & Information Centre, UNIC, with a request to upload this Circular on University Website.**

Copy to :-

1] The Director, Board of Examinations & Evaluation,Dr.BAMU,A'bad.
2] The Section Officer,[M.Sc.Unit] Examination Branch,Dr.BAMU,A'bad.
3] The Programmer [Computer Unit-1] Examinations, Dr.BAMU,A'bad.
4] The Programmer [Computer Unit-2] Examinations, Dr.BAMU,A'bad.
5] The In-charge,[E-Suvidha Kendra], Rajarshi Shahu Maharaj Pariksha Bhavan, Dr.BAMU,A'bad.
6] The Public Relation Officer, Dr.BAMU,A'bad.
7] The Record Keeper, Dr.BAMU,A'bad.

DR. BABASAHEB AMBEDKAR MARATHWADA UNIVERSITY,

AURANGABAD.

SYLLABUS

M. Sc. Chemistry

(Semester I and II)

Choice Based Credit and Grading System

Effective from: 2021-22

Dr. Pathan Mohd Arif Ail Khan
Professor,
Departmentof Chemistry
Maulana Azad College of Arts Science & Commerce,
Dr. Rafiq Zakaria Campus, Rauza Baugh, Aurangabad

28/1/21

(26 pages)

DR. BABASAHEB AMBEDKAR MARATHWADA UNIVERSITY, AURANGABAD

Curriculum under Choice Based Credit and Grading System

M.Sc. Chemistry I & II Semester

Effective from: 2021-22

DR. BABASAHEB AMBEDKAR MARATHWADA UNIVERSITY, AURANGABAD

SUBJECT OF CHEMISTRY

Syllabus Based upon Choice Based Credit and Grading System

The University conducts five specializations namely Analytical, Drug, Inorganic, Organic and Physical Chemistry.

The M. Sc. Chemistry program is divided into Four Semesters with minimum 108 credits, comprising of 18 theory courses and 12 laboratory courses including projects, seminars and tutorials.

The 18 theory courses of each specialization include:

Core Courses:	08 + 01
Specialized Courses:	07
Elective Courses:	01
Service Courses:	01

The 12 laboratory course includes:

Core Courses:	04
Specialized Courses:	04

Credit System and Cafeteria approach

- The credit system with cafeteria approach has been adopted by the University.
- Students will have to learn at least 108 credits for the award of M. Sc. Degree.
- Out of 108 credits students will have to learn at least 34 credits of core courses, 28 credits of specialized courses and 08 credits of elective/service courses.
- Along with these credits of theory courses students are required to earn at least 36 credits of laboratory courses.

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The following will be the structure of syllabus for M.Sc. Chemistry I and II semester effective from 2021-22

Semester	Paper No.	Title	Credits	Teaching Hrs/Week	Marks		Total Marks
					Internal	Exam.	
Semester-I	Compulsory	Introduction to the Indian Constitution	02	02	10	40	50
	CHE-101	Analytical Chemistry	04	04	20	80	100
	CHE-102	Inorganic Chemistry	04	04	20	80	100
	CHE-103	Organic Chemistry	04	04	20	80	100
	CHE-104	Physical Chemistry	04	04	20	80	100
Semester-II	CHE-205	Analytical Chemistry	04	04	20	80	100
	CHE-206	Inorganic Chemistry	04	04	20	80	100
	CHE-207	Organic Chemistry	04	04	20	80	100
	CHE-208	Physical Chemistry	04	04	20	80	100
I and II semester Laboratory Course	CHE-209	Laboratory Course (Analytical Chemistry)	4.5	06	NA	50	50
	CHE-210	Laboratory Course (Inorganic Chemistry)	4.5	06	NA	50	50
	CHE-211	Laboratory Course (Organic Chemistry)	4.5	06	NA	50	50
	CHE-212	Laboratory Course (Physical Chemistry)	4.5	06	NA	50	50

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The structure of uniform syllabus for M.Sc. Chemistry III semester effective from 2022-23

Semester	Paper No.	Title	Credits	Teaching Hrs/Week	Marks		Total Marks
					Internal	Exam.	
Semester-III	CHE-313 Common		04	04	20	80	100
	CHEA-314		04	04	20	80	100
	CHED-314		04	04	20	80	100
	CHEI-314		04	04	20	80	100
	CHEO-314		04	04	20	80	100
	CHEP-314		04	04	20	80	100
	CHEA-315		04	04	20	80	100
	CHED-315		04	04	20	80	100
	CHEI-315		04	04	20	80	100
	CHEO-315		04	04	20	80	100
	CHEP-315		04	04	20	80	100
	CHEA-316		04	04	20	80	100
	CHED-316		04	04	20	80	100
	CHEI-316		04	04	20	80	100
	CHEO-316		04	04	20	80	100
	CHEP-316		04	04	20	80	100
	Service Course		04	04	20	80	100

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The structure of uniform syllabus for M.Sc. Chemistry IV semester effective from 2022-23

	Paper No.	Title	Credits	Teaching Hrs/Week	Marks		Total Marks
					Internal	Exam.	
Semester IV	CHEA-417		04	04	20	80	100
	CHED-417		04	04	20	80	100
	CHEI-417		04	04	20	80	100
	CHEO-417		04	04	20	80	100
	CHEP-417		04	04	20	80	100
	CHEA-418		04	04	20	80	100
	CHED-418		04	04	20	80	100
	CHEI-418		04	04	20	80	100
	CHEO-418		04	04	20	80	100
	CHEP-418		04	04	20	80	100
	CHEA-419		04	04	20	80	100
	CHED-419		04	04	20	80	100
	CHEI-419		04	04	20	80	100
	CHEO-419		04	04	20	80	100
	CHEP-419		04	04	20	80	100
	CHEA-420A or CHEA-420B		04	04	20	80	100
	CHED-420A or CHED-420B		04	04	20	80	100
	CHEI-420A or CHEI-420B		04	04	20	80	100
	CHEO-420A or CHEO-420B		04	04	20	80	100
	CHEP-		04	04	20	80	100

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	420A or CHEP-420B						
III and IV semester Laboratory Course	Laboratory Course-I		4.5	06	NA	50	50
	Laboratory Course-II		4.5	06	NA	50	50
	Laboratory Course-III		4.5	06	NA	50	50
	Laboratory Course-IV Project (Dissertation & Presentation)		4.5	06	NA	50	50

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CHE-101 ANALYTICAL CHEMISTRY

Teaching hours: 60 **Credits: 04**

Unit- I. Statistical Treatment of analytical data:

Introduction, types of errors, significant figures, precision and accuracy, methods of expressing accuracy, methods of expressing precision, the confidence limit, tests of significance- the F test, the student T test, rejection of results - the Q test. Statistics for small data sets, linear least squares, correlation coefficient, using spreadsheets for plotting calibration curves, slope, intercept and coefficient of determination, numericals.

Unit - II. Basic Separation techniques: Distillation and Solvent and Solid Phase extraction: **12hrs**

Distillation: Fractional distillation, distillation under vacuum, theory of operation of distillation methods, some practical considerations.

Solvent and Solid Phase extraction: Phase equilibrium, the partition coefficient the distribution ratio, theory of phase contact methods, single equilibrations, repeated equilibrations, counter current distribution, practical aspects and applications - extraction of metals, extraction of molecular species, Ion pair extractions, Accelerated and microwave assisted extraction, solid phase extraction, Numericals.

Unit - III. Chromatography **12 hrs**

Introduction, basic principles and theory of chromatographic techniques, plate theory of chromatography, rate theory of chromatography, other factors in zone broadening, Development of the chromatogram - Frontal analysis, elution analysis displacement analysis, selection of chromatograph system, qualitative and quantitative analysis by chromatography.

Unit - IV. Chromatographic Systems **12 hrs**

(a) Thin layer Chromatography:

Basic principles, experimental techniques, solvent systems, plate development, detection of components, evaluation of chromatogram by different methods, applications of TLC.

(b) Liquid-Liquid partition chromatography:

Introduction, theory, solid supports, selection of stationary and mobile phases, reverse phase chromatography, choice of adsorption or partition, applications of partition chromatography.

(c) Column Chromatography:

Principle, experimental details, theory of development, column efficiency, factors affecting column efficiency, and applications.

(d) Gel permeation Chromatography:

Principle materials, gel preparation, column packing, detectors and applications.

(e) Ion Exchange Chromatography:

Ion Exchange resins, ion exchange equilibria, ion exchange capacity of resins and its determination, applications of ion exchange resins to chromatography, ion chromatography based on suppressors

Unit- V.

(a) Gas Chromatography: **12 hrs**

Introduction, principles of gas-liquid chromatography, instrumentation - Carrier gas, sample introduction system, columns, detectors, substrates, temperature control, evaluation Retention volume, resolution, branches of gas chromatography, applications, numericals.

(b) High Performance Liquid Chromatography:

Principle, instrumentation - column, column packing, mobile phone, pumping system, detector system, practical procedure, applications, HPLC adsorption and partition chromatography.

Reference Books:

1. Fundamental of Analytical Chemistry 8th Edn. Skoog, West Hollar, Couch.
2. Analytical Chemistry 6th Edn., G.D. Christian
3. Chemical Separations and Measurements, D.G. Peters, J.M. Hayes and G.M. Hieftie
4. Instrumental Method of Chemical Analysis, G.R. Chatwal & S. K. Anand.

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CHE-102 INORGANIC CHEMISTRY

Teaching hours: 60 **Credit: 04**

Unit -I Symmetry concepts **12 hrs.**

Introduction to symmetry operations, symmetry elements, point group, classifications of point groups, point group of H_2O, NH_3, CO_2, BF_3, C_2H_4, PCl_3, PCl_5, C_6H_6, $[PtCl_4]^{--}$,$[PtCl_2(NH_3)]$, $[CoCl_2(NH_3)_4]$ HCl, BeF_2, CO, $[FeF_6]$, $C_2H_2Cl_2$, o, m, & p substituted benzene molecule.

Unit -II Representation of Groups **12 hrs.**

Definition of group, properties of group, applications of point groups, Group multiplication table, matrix representation of symmetry elements. Reducible and irreducible representation, character of representation, character of matrix, Conjugate matrix, Properties of irreducible representations, Great orthogonality theorem (without proof) and it's importance, construction of character table of C_{2v} & C_{3v} point group. Mulliken symbolism rules for irreducible representations & it's applications with examples. Standard reduction formula direct product and uses.

Unit III : Reaction mechanism of transition metal complexes. **12 hr**

Classification of inorganic reactions, ligand substitution reaction and their mechanisms of octahedral complexes, acid hydrolysis, factors affecting the acid hydrolysis base hydrolysis, conjugate base mechanism, Electron transfer reaction: mechanism of inner and outer sphere electron transfer reactions in octahedral complexes.

Unit IV: Metal ligand equilibria in solution: **12 hrs**

Definition of stability constant, step wise and overall formation constant, factors affecting the stability of metal complexes with reference to the nature of metal ion and ligand. Determination of formation constant for binary complexes using pH-metric technique.

Unit V: Inorganic Systems in Biological Systems **12 Hrs.**

Essential and trace elements in biological systems and their functions, Structure and function of metalloporphyrins hemoglobin, cyclochrome and hemocyanine. Electron transfer, respiration and photosynthesis reactions. Metal deficient diseases of Fe, Zn, Cu, Mn and their therapy.

Reference books:

1. Symmetry and Spectroscopy of Molecules, K.Veera Reddy.
2. Group Theory and symmetry in Chemistry, Gurdeep Raj.Ajay Bhagi and Vinod Jain.
3. Inorganic Chemistry (Principles, Structures and Reactivity) (Fourth Edition) - J.E. Hubeey, E.A. Keitler, R.L. Keitler
4. Mechanism of Inorganic Reaction. II Edn. Fred Basolo and R.G.Pearsons.
5. Selected Topic in Inorganic Chemistry, Wahid U. Malik, G.D.Tuli and R.D.Madan.
6. Advanced Inorganic Chemistry, F.A. Cotton and Wilkinson.
7. Advanced Inorganic Chemistry, Satyaprakash, G.D.Tuli, S.K.Basu and R.D.Madan.
8. Advanced Inorganic Chemistry, Volume I and II Gurdeep Raj.
9. Concise Inorganic Chemistry, J.D. Lee.
10. A Textbook of bioinorganic chemistry, A. K. Das
11. Symmetry and group theory in Chemistry, R. Ameta

CHE-103 ORGANIC CHEMISTRY

Total teaching hrs: 60 **Credits: 04**

Unit-I: Nature of Bonding in Organic Molecules: **[12 hrs]**

Delocalized chemical bonding, conjugation, cross conjugation, resonance, hyperconjugation, bonding in fullerenes, tautomerism.

Aromaticity in benzenoid and non-benzenoid compounds, alternant and non-alternant compounds, Huckel rule, energy level of π-molecular orbitals, annulenes, aromaticity, homo-aromaticity, ψ-aromaticity, PMO approach; Bonds weaker than covalent - addition compounds, crown ether complexes and cryptands, inclusion compounds, cyclodextrins, catenanes and rotaxanes.

Unit-II: Reaction Mechanism : Structure and Reactivity **[12 hrs]**

Types of Mechanisms, Types of reactions, Thermodynamic and Kinetic requirements, Kinetic and Thermodynamic control, Hammond's postulate, Curtin-Hammett Principle, Potential energy diagrams, transition states and intermediates, methods of determining mechanisms, isotope effects, hard and soft acids and bases, Generation, structure, stability and reactivity of carbocations, Carbanions, free radicals, carbenes and Nitrenes. Effect of structure on reactivity, resonance and field effect, steric effect quantitative treatment, The Hammett equation, Linear free energy relationship, substituent and reaction constants, Taft equation.

Unit-III & IV: Stereo-chemistry: **[24 hrs]**

Elements of symmetry, chirality, Enantiomeric and diastereomeric relationships, R and S, E and Z nomenclature. Molecules with more than one chiral center, Threo and Erytheo isomers, Prochiral relationships, groups and faces, stereospecific and stereoselective reactions. Optical activity in the absence of Chiral Carbon (Biphenyls, allenes and Spiranes), Chirality due to helical shape. Methods of resolution, optical purity, stereochemistry of the compounds containing Nitrogen, Sulphur and phosphorous. Conformational analysis of cycloalkanes, Mono and disubstituted cyclohexanes, decalins, effect of conformation on reactivity.

Unit-V: Aliphatic Nucleophilic and Electrophilic Substitutions: **[12 hrs]**

Nucleophilic: The SN^2, SN^1 mixed SN^1 and SN^2 and SET mechanisms. The neighbouring group mechanism, Neighbouring group participation by π and 6-bonds, anchimeric assistance. The SN^i mechanism. Nucleophilic Substitution at an allylic aliphatic trigonal and a vinylic carbon.

Reactivity : Effect of substrate structure, attacking nucleophile, leaving group and reaction medium. Phase transfer catalysis, Ambident nucleophiles, regioselectivity. Finkelstein reaction, Appeal Delepine reactions

Electrophilic Reactions : Bimolecular mechanisms-SE^2 and SE^i. The SE^1 mechanisms. Electrophilic substitution accompanied by double bond shifts.

Reference Books:

1. Advanced Organic Chemistry, IV Edition: J. March
2. Stereochemistry of Carbon Compounds: E. L. Eliel
3. Advanced organic Chemistry, Part-A and Part-B: F. A. Carey, & R. J. Sundburg.
4. A Guide Book to Mechanism in Organic Chemistry: Peter Sykes.
5. Synthetic Organic Chemistry: H. O. House
6. Principles of Organic Synthesis: R. O. C. Norman
7. Stereochemistry of Organic Compounds: D. Nashipuri
8. Organic Chemistry: Clayden and Greeves
9. Mechanism and Structure in Organic Chemistry: E. S. Gould

CHE-104 PHYSICAL CHEMISTRY

Total Teaching Hours : 60 **Credits: 04**

Unit I: Ionic Equilibria and Biological Reactions **12 hrs.**

Exact treatment of the dissociation of week acids and bases, Dissociation constant of polyprotic acids, Statistical effects in polyprotic acids, Dissociation constant of complex ions, Logarithmic expression for pH and pOH, Calculations involving buffer solution, buffer capacity and buffer index, Salt effect and solubility product and its applications.
Thermodynamics of biochemical reactions, Binding of oxygen by myoglobin and hemoglobin, Reaction between microscopic and macroscopic dissociation constant.

Unit II: Chemical Dynamics **12 hrs.**

Collision theory, modified collision theory, weakness of the collision theory, Theory of absolute reaction rates, equilibrium hypothesis, Derivation of the rate equation, statistical mechanical derivation and thermodynamic formulation. Isotope effect on reaction rate. Primary salt effect, secondary salt effect.
Dynamics of uni-molecular reactions, Lindmann and Hinshelwood theory
Kinetics of fast reactions, study of fast reactions by flow method, relaxation method, flash photolysis and NMR method.
Reactions in solution: Reaction between ions, influence of solvent-double sphere model, single sphere model, influence of ionic strength, numericals.

Unit III: Classical Thermodynamics **12 hrs.**

Nernst heat theorem, the third law of thermodynamics, determination of absolute entropies of solids, liquids and gases. Partial molar properties : Partial molar free energy, chemical potential, partial molar volume and partial molar heat content and their significance, determination of these quantities, concept of fugacity and determination of fugacity, numerical.

Unit IV: Surface Chemistry: **12 hrs.**

Surface tension, capillary action, pressure difference across curved surface (Laplace equation) vapour pressure of droplets (Kelvin equation) Gibbs adsorption isotherm, estimation of surface area (BET equation), surface films on liquids (Electro kinetic phenomenon), catalytic activity at surfaces, numericals.
Colloidal electrolytes, Types of micelles in colloidal electrolytes, Surface active agents, Classifications of surface active agents, Micellization, critical micellar concentration, Factors affecting CMC, Thermodynamics of micellization, Mechanism of Micellization, Determinations of critical micellar concentration, Reverse micelles, Solubilization

Unit V: Electro-Chemistry **12 hrs.**

Debye-Huckel theory of strong electrolytes, Debye-Huckel-Onsager equation Testing of the equation, Debye-Falkenhagen effect, Wein effect, activity coefficient, mean ionic activity coefficient; Debye-Huckel limiting law ionic strength. Electrocapillary phenomena, and its measurements. Effect of anions, cations and molecules on electrocapillary curves. Electrocapillary properties of mercury-solution interface.
Polarography: the Ilkovic equation and its derivation, concentration polarization, instrumentation, advantages of DME, half wave potential. Applications of polarography, numericals.

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References books:

1. Chemical Kinetics - Laidler (McGraw-Hill)
2. Kinetic and Mechanism of Chemical Transformations - J. Rajaram and J.C. Curiacose (Macmillan India Ltd.)
3. Physical Chemistry - Atkins (Oxford)

4. Thermodynamics for Chemists - S. Glasstone (EWP, New Delhi)
5. Physical Chemistry - G. M. Barrow
6. Advanced Physical Chemistry - Gurdeep-Raj (Pelenum)
7. Micelles : Theoretical and Applied Aspects - V. Moroi (Plenum)
8. Text Book of Physical Chemistry - S.Glasstone (McMillan)
9. An Introduction to Electrochemistry - S. Glasstone (EWP, New Delhi)
10. Physical chemistry – Robert A .Alberty ., Robert J .Silbey
11. Statistical Thermodynamic – M. C. Gupta

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CHE-205 SPECTROSCOPIC METHODS OF ANALYSIS

Teaching hours: 60 hrs **Credits: 04**

Unit -I : General introduction of spectral methods of analysis. **12 hrs**

Characterization of electromagnetic radiations, Regions of the spectrum, Interaction of radiations with matter - absorption, emission, transmission, reflection, dispersion, polarization and representation of spectra, basic elements of practical spectroscopy, resolving power, signal to noise ratio. Uncertainty relation and natural line width, natural line broadening, intensity of spectral lines, energy levels, selection rules, components of spectrometer and their functions.

Microwave spectroscopy: Rotation of molecules, rotational spectra, diatomic molecules - rigid diatomic molecules, intensities of spectral lines, effect of isotopic substitution, non-rigid rotator, the spectrum of non-rigid rotator, polyatomic molecules, technique and instrumentation in outline, applications, numerical problems.

Unit – II : Vibrational spectroscopy **12 hrs**

Review of linear harmonic oscillator, the vibrating diatomic molecule, the simple harmonic oscillator, the anharmonic oscillator, the diatomic vibrating rotator, the vibration-rotation spectrum of carbon monoxide, breakdown of the Born-Oppenheimer approximation, the vibration of polyatomic molecules, overtones and combination frequencies, the influence of rotation on the spectra of polyatomic molecules, the influence of nuclear spin, symmetric top molecules, analysis by Infra-red technique - Group frequencies, outline of technique and instrumentation. **Raman spectroscopy:** Classical and quantum of theory of Raman effect, pure rotational, vibrational and vibrational-rotational Raman spectra, rule of mutual exclusion, overtone and combination vibrations, Rotational fine structure, outline of technique and instrumentation, applications.

Unit –III : Electronic Spectroscopy **12 hrs**

Atomic spectroscopy : Energies of atomic orbitals vector representation of momenta and vector coupling, spectra of hydrogen and alkali metal atoms.

Molecular spectroscopy: Energy levels, molecular orbitals, vibronic transitions, vibrational progression, and geometry of the excited states, Franck-Condon principle, electronic spectra of polyatomic molecules.

Photoelectron spectroscopy: Basic principles, ESCA- Introduction - ESCA - ESCA satellite peaks spectral splitting ESCA chemical shifts, instrumentation, applications, Auger electron spectroscopy (brief review)

Unit -IV. **12 hrs**

Ultraviolet- Visible Spectroscopy:

Various Electronic transitions, chromophores, Auxochromes, Bathochromic and Hypsochromic Shifts, Effect of solvent on electronic transitions, Woodward-Fieser rules for dienes, enones and aromatic compounds, Applications of U.V.

Infrared Spectroscopy

Characteristic vibrational frequencies of alkenes, alkynes, aromatic compounds, Carbonyl compounds, hydroxy compounds, amines and metal-ligand complexes. Factors affecting IR group frequencies, overtones, combination bands and Fermi resonance. Applications of IR.

Unit- V.

Nuclear Magnetic Resonance Spectroscopy **12 hrs**

Elementary Ideas, Chemical Shifts, Factors affecting chemical shifts, Spin-Spin couplings and coupling constants (J), Integration.

Interpretation of HNMR, 13CNMR, Mass IR and UV spectra of Methyl acetate, Methyl propionate, ethyl acetate, cyclohexane 1,4-dione, cyclohexane 1,3- dione, isobutyraldehyde, propionic acid, neo pentane.

References Books:

1. The Determination of Molecular Structure: P. J. Wheatley
2. Physical Chemistry : G. M. Barrow
3. Instrumental Methods of Chem. Analy. Chatwal and Anand.
4. A Text book of Phy.Chem. : A.S. Negi & S. C. Anand
5. Instrumental Methods of Chemical Analysis - Willard, Merritt, Dean & Seatle
6. Instrumental Methods of Chemical Analysis - Chatwal, Anand
7. Instrumental Methods of Chemical Analysis - B.K. Sharma
8. Instrumental Methods of Chemical Analysis -R.D. Braun
9. Analytical Chemistry : Skoog and West
10. Principles of Instrumental Analysis : Skoog and West.
11. Fundamentals of Molecular Spectroscopy : Banwell.
12. Atomic and Molecular Structure : Manas Chanda
13. Molecular Spectroscopy : B.D. Acharya
14. Molecular Spectroscopy : Dyer.
15. Organic Spectroscopy : P.S. Kalsi (6th Edition).
16. Spectroscopic Methods in Organic Chemistry : D.H. Williams and I.Fleming.
17. Spectrometric Identification of Organic Compounds : R.M. Silverstein, Morril and G.C. Bassler
18. Introduction to Spectroscopy : Pavia, Lampman and Ktiz (3rd Edition)
19. Organic Spectroscopy : William Kemp (3rd Edition).
20. Quantum Chemistry- B. K. Sen
21. Inorganic Chemistry - Atkin and Shriver.

CHE-206 INORGANIC CHEMISTRY

Teaching hours: 60 **Credit: 04**

Unit -I : Spectroscopic term symbols: **12 hrs.**

Terms, inter-electronics repulsion, spin orbit coupling, ground terms, determination of terms symbol for d^1 to d^5 Configuration / complexes, Hunds rule, microstates. Racah Parameter. Weak and stronger field approach, correlation diagram of d^1, d^2, d^8 and d^9 configuration in octahedral and tetrahedral environments, Non-crossing rule, Orgel diagram of d^1 to d^9 configuration in an octahedral and tetrahedral environments, Tanabe Sugano diagram of d^2 and d^3 configuration of an octahedral environments.

Unit II : Electronic Spectra and magnetic properties of metal complex : **12 hrs.**

Types of experimental recording of the spectra, Classification of charge transfer transitions and their mechanisms with suitable examples, Band intensities, intensity of d-d bands, intensity of change transfer bands. Interpretation of electronic spectrum of transition metal complex with suitable examples, Kongis method for calculation of Dq, B and β parameters and numericals.

Unit III : Chemistry of Metal Carbonyls **12 hrs.**

Classification; Chemistry of carbonyl group Preparation, properties, structures and bonding in - iron carbonyls, $Ni(CO)_4$, $Co_2(CO)_8$, $Mn_2(CO)_{10}$, $Cr(CO)_6$, $Mo(CO)_6$ and $W(CO)_6$, $Co_4(CO)_{12}$ and $V(CO)_6$. EAN rule applied to these carbonyls structures of mixed carbonyls of transition metals and EAN rule applied to these carbonyls. Preparations of carbonyl halides

Unit IV: Metal nitrosyl compounds **12hrs.**

Preparations and properties of Nitrosyl halides (NOX), Metal nitrosyl halides, compounds containing NO- group, Compounds containing NO+ groups, Preparation, structure and application of sodium Nitropruside. EAN and eighteen electron rules applied to: Nitrosyl compounds of Cobalt, iron and Manganese. Significance of NO for the life of living animals

Unit-V : Theories of Metal–Ligand Bonding in Complexes **12 hrs.**

Valence Bond Theory: Assumptions of VBT, VBT applied to Octahedral (inner & outer orbital) complexes, Limitations of VBT.

Crystal field Theory: Important features of CFT, Crystal field splitting of d–orbital in Octahedral, Tetrahedral, Square–planar and Tetragonal complexes, Crystal field stabilization energy, Calculation of CFSE of octahedral and tetrahedral complexes. Factors affecting the magnitude of Δo, John Teller distortion and its Applications of CFT, Limitations of CFT.

Molecular Orbital Theory: Molecular Orbitals in complexes, Principles of molecular orbital theory, construction and explanation of molecular orbital energy level diagrams of Octahedral, Tetrahedral, Square–planar complexes involving sigma– as well as Π–bonding with suitable examples

References Books:

1. Advanced Inorganic Chemistry Vol. I & Vol. II - By - Gurdeep and Raj.
2. Inorganic Chemistry (Principles, Structures and Reactivity) (Fourth Edition) By - J.E. Hubeey, E.A. Keitler, R.L. Keitler.
3. Inorganic Chemistry – Gary L Miessler, P J Fischer and Donald A Tarr.
4. Advanced Inorganic Chemistry - Vol. I - By Satyaprakash, Tuli, Basu and Madan.
5. Selected Topics in Inorganic Chemistry - By W.U. Malik, G.D. Tuli & R.D. Madan.
6. Chemistry of the Elements - By N. N. Greenwood and A. Earnshaw.
7. Inorganic electronic spectroscopy, - A.B.P. Lever.
8. Symmetry and Spectroscopy of Molecules - K. Veera Reddy.
9. Physical Chemistry through problem - Dogra and Dogra.

10.Inorganic Chemistry - Attkin and Shriver.
11.Concise Inorganic Chemistry - By J.D. Lee.
12.Element of Magnetochemistry - By A.Samal & R. L. Datta.
13.Some aspect of Crystal Field theory- T. M. Dunn, D.S.Mcclure & R. G. Person
14.Introduction to Magnotochemistry- Alan Earnshaw
15.Introduction to Ligand Field - B. N. Figgis.
16. Inorganic Chemistry – James and House
17. Prinicpals of Inorganic chemistry- B R Puri, L R Sharma and K C Puri

CHE-207 ORGANIC CHEMISTRY

Credits: 04

Teaching hrs: 60

Unit-I: Aromatic Electrophilic and Nucleophilic Substitutions: **[18 hrs]**
Electrophilic Substitutions: The arenium ion mechanism, orientation and reactivity, energy profile diagram. The ortho/para ratio, IPSO substitution, orientation in other ring system, Recapitulation of halogenations, nitration, sulphonation and Fridel Craft's reaction, diazonium coupling. **Nucleophilic Substitution:** The SN^{Ar}, SN^{1}, benzyne mechanism, Effect of substrate structure, leaving group and attacking nucleophile on reactivity.

Unit-II: Addition to Carbon –Carbon multiple bond: **[12 hrs]**
Mechanism and stereochemical aspect of addition reaction involving electrophile, nucleophile and frr radicals. Regioselectivity and chemoselectivity, orientation and reactivity, Michael addition, Sharpless asymmetric epoxidation.

Unit-III: Addition to Carbon–Hetero Multiple bond: **[12 hrs]**
Mechanism of metal hydride reduction of saturated and unsaturated carbonyl compounds, acid, ester and nitriles. Addition of Grignard reagent, Organo zinc and organo lithium reagent to carbonyl and unsaturated carbonyl compounds. Wittig reaction. Mechanisum of condensation reaction involving enolates – Aldol, Knovenagel, Claisen, Mannich, Benzoin, Perkin, Stobbe reaction, Prins reaction Hydrolysis of esters and amides.

Unit-IV: Elimination Reactions: **[12 hrs]**
The E_1, E_2, and E_{1CB} mechanism, orientation of double bond. Reactivity: effect of substrate structure, attacking base, the leaving group and the medium, pyrolytic elimination.

Unit-V: Rearrangements: **[06 hrs]**
General mechanistic consideration, nature of migration, migratory aptitude, memory effect, pinacole. pinacolone, Benzil–Bezilic acid, Beckmann, Hoffman and Fries rearrangements.

Reference Books:

1. Advanced Organic Chemistry, IV Edition: J. March
2. Advanced organic Chemistry, Part-A and Part-B: F. A. Carey, & R. J. Sundburg.
3. A Guide Book to Mechanisum in Organic Chemistry: Peter Sykes.
4. Synthetic Organic Chemistry: H. O. House
5. Principles of Organic Synthesis: R. O. C. Norman
6. Organic Chemistry: Clayden and Greeves
7. Mechanism and Structure in Organic Chemistry: E. S. Gould

CHE-208 PHYSICAL CHEMISTRY

Teaching Hours: 60 **Credits: 04**

Unit - I: Quantum Chemistry: I **12 hrs.**

The Schrodinger equation, particle in a one dimensional box, Eigen values and Eigen functions, operators, properties of quantum mechanical operators, Hermitian, Linear, Ladder, Hamiltonian and angular momentum operators.

Particle in three dimensional box, harmonic oscillator, rigid rotator and numericals.

Unit - II: Quantum Chemistry: II **12 hrs.**

Term symbols and selection rules, spin-orbital coupling, The variation theorem, non-degenerate perturbation theory and its applications. Huckel molecular orbital theory of conjugated systems, application to ethylene, butadiene, cyclopropenyl radical , cyclobutadiene and benzene, numericals.

Unit –III : Phase Rule: **12 hrs.**

Recapitulation of phase rule and terms involved in it, one component system, two component systems (solid-solid, solid-liquid and liquid-liquid), reduced phase rule, three component systems, partially miscible three liquid systems : one partially miscible pair, two partially miscible pairs, three partially miscible pairs, systems composed of two solids and a liquid : crystallization of pure components only, formation of binary compounds, formation of ternary compounds, formation of solid solutions, partial miscibility of solid phases, numericals.

Unit –IV: Crystallography **12 hrs**

Classification of solids on the basis of shapes, and bonding, crystal lattice and unit cell, laws of crystallography crystal symmetry, symmetry elements, lattice planes and their designations, liquid crystals.

Principle of crystal structure. close packing of atoms, packing of equal sized spheres in HCP, CCP, BCC structures. packing in ionic solids, ionic radius, radius ratio rule, (3, 4, 6, 8 coordinate structures). octahedral and tetrahedral voids, isomorphism and polymorphism, numericals.

Unit –V : Photochemistry **12 hrs.**

Absorption of light and nature of absorption spectra, electronic transitions. photo-dissociation and pre-dissociation. photo-oxidation, photo-reduction and photo-dimerization. photo-physical phenomenon. Jablonski diagram. photo-physical pathways of molecular de-excitation, difference between delayed fluorescence and phosphorescence, Stern-Volmer equation, deviations from Stern-Volmer equation, concentration dependence of quenching and eximer formation, quenching of fluorescence formation of excimer and exciples.

References Books:

1.Quantum Chemistry : Ira N. Levine
2.Quantum Chemistry : R.K. Prasad
3.Quantum Chemistry : B.K. Sen
4.Principles of Physical Chemistry : Puri, Sharma, Pathania
5.Advanced Physical Chemistry : Gurdeep - Raj, Plenum.
6.Physical Chemistry : Maron and Prutton
7.Introduction to Molecular Photo-chemistry : C.H.J. Wells
8.Fundamentals of Photo-chemistry : Rohatgi-Mukherjee.
9.Photo-chemistry : J.G. Calvert & J.N. Pitts.
10.Ptoho-luminiscence of solutions : C.A. Parker.
11.Photo-chemistry : A. Singh and R. Singh
12Atkin's Physical Chemistry : Peter Atkins
13.Solid State Chemistry : D.K. Chakraborti

14.Solid State Chemistry and its applications : A.R. West.
15.The Determination of Molecular Structure : P.J. Wheatley.
16.Solid State Chemistry : N.B. Hannary.
17.Principles of Solid State : H.V. Keer.
18.Physical Chemistry : G.K. Vemulapalli.

CHE -209 **LABORATORY COURSE (Analytical Chemistry)**

Laboratory work hours: 135 hrs. **Credit : 4.50**

List of Experiments

1. Determination of saponification value of an oil sample.
2. Determination of active chlorine in the given sample of bleaching powder.
3. Determination of ion exchange capacity of given ion-exchange resin.
4. Determination of Mg^{++} from given sample of talcum powder.
5. Determination of aspirin in the given tablet.
6. Determination of molality of given unknown solution by Volhard method.
7. Determination of Hardness of the water sample.
8. Determination of pKa value of given substituted Benzoic acid
9. Determination of chemical oxygen demand (COD) of the given water sample.
10. Determination of Cu^{2+} ion in the given solution spectrophotometrically.
11. Determination of dichromate & permanganate ion simultaneously in the given sample spectrophotometrically.
12. Determine the molecular weight of a given polymer by turbidimetry
13. Determine the concentration of sulphuric acid, acetic acid & copper sulphate in the givensolution by conductometric titration method.
14. Estimation of Na/ K/ Li/ Ca by Flame photometry
15. Determination of Phosphoric acid cola averages by pH meter
16. Estimation of Vitamin C by 2,6 dichloro-indophenol method
17. Assay sulpha drugs.

Scheme of marking:

1) Experiment I : 20 Marks
2) Experiment –II : 20 Marks
3) Record book & Viva : 10 Marks

CHE -210 LABORATORY COURSE (INORGANIC)

Credit: 4.5

Laboratory work hours: 135 hrs.

List of experiments

I) Semi micro Qualitative Inorganic analysis. **06 mixture**

Identification of three acidic and three basic radicals including one rare earth metal from given mixture by using semi micro qualitative analysis method.

Note : Each mixture should contain different rare earth elements

II) A. Separation and estimation of metal ions from the following binary mixture solutions : **Any[04]**

1. Copper- Nickel	2. Copper- Iron	3. Nickel- Zinc
4. Iron- Magnesium	5. Copper- Barium	6. Iron –Aluminium

III).Synthesize, characterization and estimation of metal ion from the metal complexes.

Any [05]

1. $Ti(C_9H_8NO)_2.2H_2O$	2.$VO(acac)_2$	3.Cis-$K[Cr(C_2O_4)_2(H_2O)_2]$
4. $[Mn(acac)_3]$	5.$K_3[Fe(C_2O_4)_3]$	6.$[Co(II)(Py)_2Cl_2]$
7. $[Co(III)(NH_3)_6]Cl_3$	8.$[Co(III)(NO_2)(NH_3)_5]Cl_2$	9.[Ni $(NH_3)_6]Cl_2$

Note : i).Synthesis should be carried out using (0.02 to 0.06 mole) of the starting materials.
ii). Practical Yield, % yield, theoretical yield and percentage of metal ion content should be recorded

Note : Student will not be allowed for practical examination if his/her record book is not completed and certified.

Reference Book:

1. A Text book of Micro and Semi micro Qualitative Inorganic Analysis ; IV edn, A. I. Vogel
2. A Text book of Quantitative Inorganic Analysis; A. I. Vogel
3. Practical Inorganic Chemistry; Pass Geoffrey and Haydn Sutcliffe.
4. Advanced Practical Inorganic Chemistry; Gurudeep Raj;.
5. Vogel's Qualitative Inorganic Analysis, D. Svehla, VII Edn.Orient Longman Ltd.

Scheme of Marking:

I Semi micro Inorganic analysis: **25 Marks**

i). Preliminary Test :	[Max. Marks:04]
ii). Acidic radicals with C.T. :	[Max. Marks:09]
iii). Group identification of basic radicals	[Max. Marks:03]
iv)Basic radicals alongwith spot test & C.T .	[Max. Marks:09]

Note : At least one spot test for each radical should be performed and be reported

II) Separation & Estimations: **15 Marks**

ii). Estimation of metals by gravimetric / volumetric method ,
observation table, : [Max. Marks:10]
iv). Correct calculation & reporting results : [Max. Marks:05]
III) Synthesis & estimation of Metal Complexes **15 Marks**
ii). Yield of complexes : [Max. marks: 10]
iii). Estimation of metal percentage by gravimetric / volumetric method,
observation table, Correct calculation: & reporting of results : [Max. Marks: 05]

IV) Record Book & Viva : [Max. Marks 10]

CHE -211 LABORATORY COURSE (ORGANIC)

Laboratory work hours: 135 hrs. **Credit: 4.5**

List of experiments:

1) Qualitative Organic Analysis: **[30 Marks]**

Separation, purification and identification of binary mixtures.

The separation should be carried out using ether/ dichloromethane.

The two components may be solid-solid, solid-liquid or liquid-liquid (non-volatile).

The water soluble solid/liquid should also be given.

Student should submit the purified samples of the separated compounds and prepare a suitable derivative of the two compounds separated out.

Note : Analysis of at least ten mixtures should be carried out.

1) **Single Stage Preparations**: **[15 Marks]**
i) Benzaldehyde to cinnamic acid (Perkin Reaction).
ii) o-Iodo or o-chlorobenzoic acid from Anthranilic Acid.
iii) β-benzoyl propionic acid from succinic anhydride and benzene (Friedel-Craft reaction)

iv) p-nitro acetanilide from acetanilide.
v) p-nitrobromobenzene from bromobenzene.
vi) Dibenzal acetone from Benzaldehyde

vii) Salicylaldehyde from phenol (Reimer-Tiemann Reaction).

Note:

i) The preparations should be carried out using (0.02 to 0.05 mole) of the starting material.

ii) The yield, melting point and TLC of the recrystallized product should be recorded.

iii) The sample of the purified product and TLC plate should be submitted for inspection.

Note : Student will not be allowed for practical examination if his/her record book is not completed and certified.

Scheme of Marking:

1. Qualitative Organic Analysis	**25** (Max. Marks)
	Marks
Type of the mixture	06
i). Analysis of the individual components:	
ii). Detection of Elements	02
iii). Detection of functional groups	02
iv). Determination of MP/BP	02
v). Preparation of the derivative	02
&	
Identification (Spotting)	03
2. Preparation	**10** (Max. Marks)
i). Yield of the recrystallized product	05
ii). MP of the recrystallized product	05
iii). TLC of the recrystallized product	05
3) Record Book + Viva voce	**10** (Max. Marks)

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University, Aurangabad

CHE -212 LABORATORY COURSE (PHYSICAL)

Laboratory work hours: 135 hrs. **Credit: 4.5**

List of experiments

A. Instrumentation.

1.Determination of strengths of halides in a mixture potentiometrically.
2.Determination of dissociation constants of phosphoric acid potentiometrically.
3.Determination of dissociation constants of weak acid potentiometrically.
4.Determination of acidic and basic dissociation constants of an amino acid and its isoelectric point.
5.Determination of the strength of strong and weak acid in a given mixture conductometrically.
6.Determination of solubility and solubility product of sparingly soluble salt $BaSO_4$.
7.Study of kinetics of inversion of cane sugar.
8.Determination of equilibrium quotient for the formation of monothiocynato iron (III) complex.
9.Determine the indicator constant of given indicator by colorimetric measurements.
10.Determine the pK_1 and pK_2 value of phosphoric acid by pH metry.
11.To study the kinetics of mutarotation of glucose/fructose potentiometrically.

B. Non-Instrumentation.

1. Determine the molecular refraction of methyl acetate, ethyl acetate, n-hexane and carbon tetrachloride and calculate the refraction of CH_2, C, H and O atoms.
2. To study the effect of surfactants (sodium chloride) on surface tension of given liquid.
3. To determine the radius of molecule by viscosity measurements.
4. To study the adsorption of acetic acid from aqueous solution by activated charcoal and examine the validity of Freundlich and langmuir's isotherm.
5. To construct the phase diagram for three component system (chloroform-acetic acid-water).
6. Determine the solubility of benzoic acid in water at different temperature and hence its heat of solution.
7. Determine the velocity constant of hydrolysis of ester.
8. To study auto catalysis reaction between potassium permanganate and oxalic acid.
9. Determine the rate constant of the reaction between potassium persulphate and potassium iodide having equal/unequal concentration of the reacting species.
10. Determine the formula of the complex formed between Cu(II) and ammonia by distribution method.
11. To study the variation of viscosity with the composition of mixtures (ethanol-water-HNO_3-chloroform) and to determine the formation of complex between two liquids.

Note : Student will not be allowed for practical examination if his/her record book is not completed and certified.

Scheme of marking:

1) Experiment I(Instrumentation) : 20 (Max. Marks)
2) Experiment –II (Non-Instrumentation) : 20 (Max. Marks)
3) Record book & Viva : 10 (Max. Marks)

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Theory Question Paper Pattern

Question NO 1 is compulsory
Solve any five questions from question number 2 and 8

Question No. 1

A. Multiple-choice questions (10 questions of 1 Mark each) Max. Marks 10

B. Short Answer type question (Five questions of 2 marks each) Max. Marks 10

Question No. 2 (Two or three bit) Max. Marks 12

a.
b.
c.

Question No. 3 (Two or three bit) Max. Marks 12

a.
b.
c.

Question No. 4 (Two or three bit) Max. Marks 12

a.
b.
c.

Question No. 5 (Two or three bit) Max. Marks 12

a.
b.
c.

Question No. 6 (Two or three bit) Max. Marks 12

a.
b.
c.

Question No. 7 (Two or three bit) Max. Marks 12

a.
b.
c.

Question No. 8 (Two or three bit) Max. Marks 12

a.
b.
c.

Dr. Pathan Mohd Arif Ail Khan
Professor,
Department of Chemistry
Maulana Azad College of Arts Science & Commerce,
Dr. Rafiq Zakaria Campus, Rauza Baugh, Aurangabad

Dean
Faculty of Science & Technology
Dr. Babasaheb Ambedkar Marathwada
University, Aurangabad

DR. BABASAHEB AMBEDKAR MARATHWADA UNIVERSITY,

AURANGABAD.

SYLLABUS

B. Sc. Chemistry

(Semester I)

(Bridge Course)

Choice Based Credit and Grading System

Effective from: 2021-22

Dr. Pathan Mohd Arif Ail Khan
Professor,
Departmentof Chemistry
Maulana Azad College of Arts Science & Commerce,
Dr. Rafiq Zakaria Campus, Rauza Bagh, Aurangabad

(2 Pages)

B.Sc. Chemistry

Bridge Course Syllabus

2 Credits 30 Hrs

Analytical Chemistry (Unit-1) **06 Hrs**

Acids, bases, salts, Types of analysis, Types of titrations, volumetric apparatus, calibration of pipette and burette. Indicators used in pH - titrations, oxidizing agents used in titrations, redox titration, errors and their impact on analysis

Mathematical Concepts (Unit-II) **08 Hrs**

Logarithmic relations, curve sketching, linear graphs and calculation of slopes, differentiation of functions like k^x e^x , x^n , sin x, log x; maxima and minima, partial differentiation. Atomic orbital's, Quantum numbers, Heisenberg uncertainty principle, shapes of s, p, d orbital's, Aufbau and Pauli exclusion principles. Hund's multiplicity rule. Electronic configurations of the elements

Chemical Bonding (Unit-III) **08 Hrs**

Atomic and Ionic radii, Ionization Energy, Electron affinity and Electro negativity, Trends in periodic table and application in predicting and explaining the chemical behavior, Covalent Bond - Valence theory, Ionic Bonds – Definitions, Hydrogen bonding, Van-der-Waals forces, Metallic bond. Localized and delocalized chemical bond; resonance, hyper conjugation, inductive effect, hydrogen bonding, conjugative effect, steric effect

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Organic Chemistry (Unit-IV) **08 Hrs**

Classification of organic compound, functional groups, Nomenclature of organic compounds. Concept of Isomerism, Optical Isomerism - elements of symmetry, molecular chirality, enantiomers, stereogenic centre, optical activity, D and L and R and S systems of nomenclature, Geometric Isomerism. Homolytic and heterolytic bond breaking. Types of reagents electrophiles and nucleophiles, Types of organic reactions, Reactive intermediates - carbocations, carbanions, free radicals.

Dr. Pathan Mohd Arif Ail Khan
Professor,
Departmentof Chemistry
Maulana Azad College of Arts Science & Commerce,
Dr. Rafiq Zakaria Campus, Rauza Baugh, Aurangabad

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SYLLABUS

M. Sc. Chemistry

(Semester I)

(Bridge Course)

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Dr. Pathan Mohd Arif Ail Khan
Professor,
Department of Chemistry
Maulana Azad College of Arts Science & Commerce,
Dr. Rafiq Zakaria Campus, Rauza Baugh, Aurangabad

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Dr. Babasaheb Ambedkar Marathwada
University, Aurangabad

(3 pages)

M.Sc. Chemistry

Bridge Course, Syllabus

2 Credits **30 Hrs**

Spectroscopy (Unit I) 06 Hours

UV-Visible Spectroscopy: Types of electronic transitions, λmax, Chromophores and Auxochromes, Bathochromic and Hypso-chromic shifts, Intensity of absorption;

IR Spectroscopy: Fundamental and non-fundamental molecular vibrations; IR absorption positions of O, N and S containing functional group, Fingerprint region and its significance;

NMR Spectroscopy: Basic principles of Proton Magnetic Resonance, chemical shift and factors influencing it; Spin-spin coupling and coupling constant; Anisotropic effects in alkene, alkyne, aldehydes and aromatics; Interpretation of NMR spectra of simple compounds.

Mass Spectroscopy-Basic principle, Fragmentation, Rule thirteen, Nitrogen rule, Determination of m/e ratio.

Basics of Organic Chemistry (Unit-II) 08 Hours

Electronic Displacements: Inductive, electrometric, resonance and mesomeric effects, hyper conjugation and their applications; Dipole moment; Organic acids and bases. Homolytic and Heterolytic fission. Curly arrow rules; Electrophiles and Nucleophiles; Nucleophilicity and basicity; Types, shape and their relative stability of carbocations, carbanions, free radicals and carbenes.

Reaction Mechanisms Introduction to types of organic reactions and their mechanism: Addition, Elimination and Substitution reactions Elimination reactions, E1, E2 & E1cb reactions. Saytzeff and Hofmann eliminations. Electrophilic additions their mechanisms (Markownikov/Anti-Markownikov addition. Electrophilic substitutions in aromatic system and aliphatic system. Nucleophilic substitutions SN1, SN2 and SNi mechanism. Role of solvent and stereochemistry in reaction mechanism. Lewis structure, sawhorse, Fischer's, Newman's formulae and their interconversions

Inorganic Chemistry (Unit-III) 08 Hours

Periodicity of elements: Periodicity of elements: s, p, d, f block elements, the long form of periodic table. Discussion of the following properties of the elements, with reference to s & p-

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block. (a) Effective nuclear charge, shielding or screening effect, Slater rules, variation of effective nuclear charge in periodic table. (b) Atomic radii (van der Waals) (c) Electronegativity, Paulings / Mullikens electronegativity scales. Ionic bond: General characteristics, Covalent bond: Lewis structure.

Pauli's Exclusion Principle, Hunds rule of maximum multiplicity, Aufbau's principle and its limitations

Physical Chemistry (Unit-IV) 08 Hours

Atomic structure: Bohr's theory, its limitations and atomic spectrum of hydrogen atom. Wave mechanics: de Broglie equation, Heisenberg's Uncertainty Principle and its significance, Schrodinger's wave equation,

Chemicals Kinetics: Chemical Kinetics and its scope, rate of reaction, factors influencing the rate of reaction - concentration, temperature, pressure, solvent, light, catalyst concentration dependence of rates. Derivation of rate law and characteristics of simple chemical reactions - zero order, first order, second order, Pseudo order, half life.

Dr. Pathan Mohd Arif Ali Khan
Professor,
Departmentof Chemistry
Maulana Azad College of Arts Science & Commerce,
Dr. Rafiq Zakaria Campus, Rauza Baugh, Aurangabad

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